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STUDY OF CARBON CHEMISTRY

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INTRODUCTION: Carbon chemistry is so important that it has a whole branch of chemistry entirely devoted to it - organic chemistry. The number of compounds that contain carbon vastly exceeds all other compounds combined. This course explores how to teach about the special nature of carbon, some of the important classes of compounds it forms, and their most important reactions. It will help you to help your students understand how the complexity of these compounds leads to variation in molecular structures and spatial arrangements of atoms – the topic of isomerism – and how we can use a variety of methods to find out more about these structures. Topics that students traditionally find difficult, such as reaction mechanisms, are unpicked in detail, and resources are analysed to give concrete approaches for helping students in in the classroom.

After working through this course you will be able to:

- Confidently teach about the key aspects of carbon chemistry.
- Help students understand and represent structures of carbon compounds;
- confidently teach students about functional groups, their main specific reactions and the principles that govern reactions involving carbon compounds;
- help students explain the factors that lead to variation in the physical properties of carbon compounds;
- •confidently teach about organic synthesis and methods for determining the structures of carbon compounds.
- •Carbon (C) appears in the second row of the periodc table and has four bonding electrons in its valence shell see our Periodic Table module for more information.

Similar to other non-metals, carbon needs eight electrons to satisfy its valence shell. Carbon therefore forms four bonds with other atoms each bond consisting of one of carbon's electrons and one of the bonding atom's electrons. Every valence electron participates in bonding; thus, a carbon atom's bonds will be distributed evenly over the atom's surface.

The second class of simple hydrocarbons, the alkenes, consists of molecules that contain at least one double-bonded carbon pair. Alkenes follow the same naming convention used for alkanes. A prefix (to describe the number of carbon atoms) is combined with the ending "ene" to denote an alkene. Ethene, for example is the two-carbon molecule that contains one double bond. The chemical formula for the simple alkenes follows the expression CnH2n. Because one of the carbon pairs is double bonded, simple alkenes have two fewer hydrogen atoms than alkanes.

In addition to carbon and hydrogen, hydrocarbons can also contain other elements. In fact, many common groups of atoms can occur within organic molecules, these groups of atoms are called functional groups. One good example is the hydroxyl functional group. A few professional and/or vital evaluations have tested international warming on the subject of the weight problems epidemic. However, none of them performed a complete systematic overview of the applicable literature, and thus, their findings had been compromised with the aid of using a loss of representativeness and challenge to take a look.

Factors affecting the properties of edible film prepared from mung bean proteins Influence of Plasticizers on the prosperities of edible film from Mung Bean Protein. He proposed the study on the results of the Edible film formed from Mung bean protein and analyzed the property affecting the edible film. Oxygen permeability and mechanical properties of banana films formed from Banana flesh. The Study shows the Methods on preparation of Edible films from Banana and further discussed about the Oxygen permeability, thickness, sealability and the mechanical properties of the film formed with clear analysis and results.